**Stoichiometry / Limiting Reagent Quiz Review: Honors**

Given the unbalanced equation below, solve the following questions:

**\_\_\_\_ NaOH + \_\_\_\_ K3PO4 → \_\_\_\_ Na3PO4 + \_\_\_\_ KOH**

1. How many grams of sodium hydroxide will it take to make 55.0 grams of potassium hydroxide, given an excess of potassium phosphate?
2. If I wanted to do this reaction and had only 22 grams of sodium hydroxide and 57 grams of potassium phosphate:
3. How many grams of potassium hydroxide could I make?
4. What is the limiting reagent?
5. How much of the excess reagent will remain?
6. If you perform the reaction and make 22.0 grams of KOH, what was your percent yield? Is this answer reasonable?
7. Define the following terms:

* limiting reagent:
* stoichiometry:

**Stoichiometry / Limiting Reagent Quiz Review: Honors**

Given the unbalanced equation below, solve the following questions:

1. **NaOH + 1 K3PO4 → 1 Na3PO4 + 3 KOH**
2. How many grams of sodium hydroxide will it take to make 55.0 grams of potassium hydroxide, given an excess of potassium phosphate?
   1. **grams**
3. If I wanted to do this reaction and had only 22 grams of sodium hydroxide and 57 grams of potassium phosphate:
4. How many grams of potassium hydroxide could I make? **30.86 g**
5. What is the limiting reagent? **NaOH**
6. How much of the excess reagent will remain? **18.08 g K3PO4**
7. If you perform the reaction and make 22.0 grams of KOH, what was your percent yield? Is this answer reasonable? **71%, yes**
8. Define the following terms:

* limiting reagent:
* stoichiometry: